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the empirical

formula of a

compound given

that the compound

is found to be 47.9

percent zinc and

52.1 percent

chlorine by mass

20 this molarity

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_____ 1) The value
of Ksp of AgCl is
 1.8×10^{-10} . What

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would be the molar concentration of Ag^+ and Cl^- in pure water placed in

contact with solid

$\text{AgCl}(s)$? $K_{sp} = [\text{Ag}^+][\text{Cl}^-] = 1.8 \times 10^{-10}$

$= [x][x]$ $x = 0.000013416 \text{ M}$ or $1.3416 \times 10^{-5} \text{ M}$

2) The molar solubility of PbCl_2 ...

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many protons and
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element that has
present in a
vanadium atom?

atoms that contain
25 protons? 2. How
many protons and
electrons are 9.

What is the name
of the element that
has present in a
nitrogen atom?

atoms that contain
82 protons? 3. How
many protons and

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0.710 kg of sodium carbonate (washing soda), Na_2CO_3 ,

in 10.0 kg of

water—a saturated

solution at 0°C ;

125 g of NH_4NO_3

in 275 g of

water—a mixture

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used to make an instant ice pack; 25 g of Cl_2 in 125 g of dichloromethane, CH_2Cl_2 ; 0.372 g of histamine, $\text{C}_5\text{H}_9\text{N}$, in 125 g of chloroform, CHCl_3

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