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The sum of the protons and neutrons in the nucleus of an atom. $MASS\# = \# \text{ protons(atomic number)} + \# \text{neutrons}$. $\#NEUTRONS = Mass \# - \#Protons(atomic number)$ $\#Protons = Mass \# - \#Neutrons$. ISOTOPES: Are atoms of the same element that have different numbers of. neutrons and different mass numbers.

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2) All atoms of a given element are identical in size, mass, and chemical properties. 3) Atoms of a specific element different from those of another element. 4) Different atoms combine in simple whole-number ratios to form compounds. 5) A chemical reaction, atoms are separated combined or rearranged.

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