

Chemical Reactions Study Guide Answers

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Chemical Reactions Study Guide Answers

STUDY GUIDE Section 9.2 Classifying Chemical Reactions In your textbook, read about synthesis, combustion, decomposition, and replacement reactions. Assume that Q, T, X, and Z are symbols for elements. Match each equation in Column A with the reaction type it represents in Column B. Column A 2. $Q + Z \rightarrow QZ$ Answer the following questions. a. b. c. d.

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In a chemical reaction, chemical (blank) are formed or broken exothermic; product During an (blank) reaction, more energy is required to form the bonds in the (blank) than needed to break the bonds of the product are formed.

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Concentration increase, less surface area, higher temperature, and a catalyst. Four ways a chemical reaction has taken place. Color change, the formation of a precipitate, gas formed, temperature change. 3 examples of a physical change.

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single replacement reaction. answer. A chemical change in which one element replaces a second element in a compound. question. double replacement reaction. answer. A chemical change involving an exchange of positive ions between two compounds. question. combustion reaction.

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Study Guide for Content Mastery Answer Key Chemistry: Matter and Change . o x T a ; 0 o u o z o z o S 0 r m 0 0 o o O o O o o 0 o 0 0 E o 0 0 o 0 o u u o o 0 o o 0 o o o 0 2 0 o N x N u N o m g g 0 0 o O o N O N O m o z N 9 o r r N r z

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Chapter 9 Study Guide – McGraw Hill Textbook. Used to show what a chemical reaction yields (same as =); used to separate the reactants from the product. Left side items are reactants and right side items are the product. (g) = gas; (aq) = water; (l) = liquid. these symbols used to denote the state of the reactants and products in an accurate balanced chemical equation.

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1. State the term for the amount of energy that is needed for a chemical reaction to occur. 2. Summarize the relationship between an enzyme and a substrate. enc 3. Classify which of the compounds in the reaction below are reactants and which are products. $\text{PbO}_2 + 4\text{HCl} \rightarrow \text{PbCl}_2 + 2\text{H}_2\text{O}$ 4. Compare endothermic reactions and exothermic reactions. f 5.

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View Test Prep - Chemical reactions study guide from CHEM 1100 at University Of Georgia. Unit 6 Study Guide for Chemical Reactions Please answer the following questions: 1. What is a chemical

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: Answer the questions and complete statements from viewing the video. Everything is made of chemicals. When do chemical reactions happen? When the electrons hook together. Water is . two. parts “H” and . one. part “O”. What makes iron rust and our blood red? Oxygen . Car and motorcycle engines are controlled by what chemical reaction?

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84 Study Guide for An Introduction to Chemistry Section Goals and Introductions Section 7.1 Energy Goals To introduce the terms energy, kinetic energy, and potential energy. To

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introduce the Law of Conservation of Energy. To describe the relationships between stability, capacity to do work, and potential energy. To explain why breaking chemical bonds requires energy and why the formation of

Chapter 7 Energy and Chemical Reactions

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For most chemical reactions in this book, solids are labeled (s), liquids (l), and gases (g). The numerical coefficients in front of the chemical formulas express the moles of each compound or element. The preceding reaction can be interpreted in terms of moles or masses. (See Table 1.)

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When a chemical reaction occurs, the original substances put together, called reactants, lose their chemical properties and become different substances called products with a different set of chemical properties. Reactions where energy is released are called exothermic reactions. When energy is absorbed, it is called an endothermic reaction.

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