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Series of cyclodextrin-closododecaborate dianion complexes, $-\text{CD} \cdot \text{B}_{12}\text{X}_{12}^{2-}$ ($\text{X} = \text{H}, \text{F}$), ; $\text{X} = \text{H}, \text{F}$). The measured vertical/adiabatic detachment energies (VDE

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photoelectrons emitted from solids, gases, or liquids by the photoelectric effect. Depending on the source of ionization energy, PES can be divided accordingly into Ultraviolet Photoelectron Spectroscopy (UPS) and X-ray Photoelectron Spectroscopy

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(XPS). The source of radiation for UPS is a noble gas discharge lamp, usually a He discharge lamp.

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(UPS) and X-ray photoelectron spectroscopy (XPS). XPS is also known under its former name of electron spectroscopy for chemical analysis (ESCA). UPS focuses on ionization of valence electrons while XPS is able to go a step further and ionize core electrons

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Abstract. Modern opto-electronics technologies are built on the basis of fundamental knowledge about

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electronic properties of semiconductor materials, which can be attained through accurate analyses made on single crystals of the materials as standard samples. In the cases of organic semiconductor materials, although technical difficulties have

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obstructed direct measurements of electronic properties in the organic semiconductor single crystals by photoelectron spectroscopy, recent advances ...

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Concepts illustrated: vibrational structure and Franck – Condon principle; adiabatic and vertical ionization energies; Koopmans's theorem; link between photoelectron spectra and molecular orbital diagrams; Morse potentials.. Carbon monoxide was

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one of the first molecules studied by ultraviolet photoelectron spectroscopy [1]. A typical HeI spectrum is shown in Figure 13.1.

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In the analytical technique of photoelectron spectroscopy (PES), a sample is ionized using high-energy radiation, and the kinetic energies of the ejected electrons (called photoelectrons) are measured. From this, we can determine the binding or ionization

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Photoelectron spectroscopy is based on Einstein's photoelectric effect. 1 A photon can ionize an electron from a molecule if the photon has an energy greater than the energy holding the electron in

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the molecule. Any photon energy in excess of that needed for ionization is carried by the outgoing electron in the form of kinetic energy.

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