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$H P_{vap} = \exp \left(\frac{14,790}{T - K} \right)$ and by the ideal gas law ISolutions to Chemical and Engineering Thermodynamics, 3e N V RT x 18 mol x 8.314 x 10⁻⁵ (bar m³ mol K) TP= = VV 1 x 10⁻³ - (10⁻³ x) 106 volume taken up by liquid Equating P and P_{vap} we have x 18 x 8.314 x 10⁻⁵ T F 5432.8 x l.

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